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Question 8. The reaction of cyanamide,  $\text{NH}_2\text{CN}$  (s) with dioxygen was carried out in a bomb calorimeter and  $\Delta U$  was found to be  $-742.7 \text{ kJ mol}^{-1}$  at  $298 \text{ K}$ . Calculate the enthalpy change for the reaction at  $298 \text{ K}$ .  $\text{NH}_2\text{CN} (\text{s}) + 3/2 \text{O}_2 (\text{g}) \rightarrow \text{N}_2 (\text{g}) + \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{l})$  Answer:  $\Delta U = -742.7 \text{ kJ mol}^{-1}$ ;  $\Delta n_g = 2 - 3/2 = +1/2 \text{ mol}$ .

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NCERT Solutions Class 11 Chemistry Chemistry Lab Manual Chemistry Sample Papers. NCERT TEXTBOOK QUESTIONS SOLVED. Question 1. Explain the formation of a chemical bond. Answer: According to Kossel and Lewis, atoms combine together in order to complete their respective octets so as to acquire the stable inert gas configuration. This can occur in two ways; by transfer of one or more electrons ...

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NCERT Solutions for Class 11 Chemistry Chapter 7 HOTS Questions Question 1. For the equilibrium  $2 \text{NOCl} (\text{g}) \rightleftharpoons 2 \text{NO} (\text{g}) + \text{Cl}_2 (\text{g})$  the value of the equilibrium constant  $K_c$  is  $3.75 \times 10^{-6}$  at  $1069 \text{ K}$ . Calculate the  $K_p$  for the reaction at this temperature?

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NCERT Solutions for Class 11 Chemistry: Chapter 2 (Structure of Atom) are provided on this page as a free source of educational content for Class 11 students. These solutions aim to provide students with comprehensive answers to all questions asked in Chapter 2 of the NCERT Class 11 Chemistry textbook.

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